



## General Instruction



- Only write with pen. Your calculator should be non-programmable.
- This theoretical exam booklet contains **71 pages**, not including general instructions.
- This examination has **9 problems**.
- You will have **5 hours** to solve the exam.
- **Begin** only when the **START** command is given.
- All results must be written in the appropriate boxes in pen in the designated areas on the **answer sheets**. Use the back of the exam sheets if you need scratch paper. Remember that answers written outside the answer boxes will not be graded.
- Write relevant calculations in the appropriate boxes when necessary. Full marks will be given for correct answers only when your work is shown.
- The invigilator will announce a **30-minute** warning before the Stop command.
- You **must stop** working when the **STOP** command is given. Failure to stop writing will lead to the nullification of your exam.
- The official English version of this examination is available on request only for clarification.
- You are not allowed to leave your working place without permission. If you need any assistance (broken calculator, need to visit a restroom, etc), raise hand and wait until an invigilator arrives.

**GOOD LUCK!**

**Problems & Grading Information**

Problem No	Title	Total Score	% of Total Score
1	Two Beauties of Turkey: the Van Cat and the Ankara Cat	24	8
2	A Tale of a Reactive Intermediate	77	10
3	(±)-Coerulescine	51	8
4	Symmetry Does Matter!	66	10
5	Konya, Carrot, Beta-Carotene, Vitamin-A, Immune System, Vision	100	14
6	Thermodynamics through an Interstellar Journey	80	12
7	Phthalocyanines	85	12
8	Boron Compounds and Hydrogen Storage	58	14
9	Quantification of Heavy Metal Ions	100	12
	<b>TOTAL</b>	<b>641</b>	<b>100</b>



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## Physical Constants and Equations

Avogadro's number	$N_A = 6.0221 \times 10^{23} \text{ mol}^{-1}$
Boltzmann constant	$k_B = 1.3807 \times 10^{-23} \text{ J K}^{-1}$
Universal gas constant	$R = 8.3145 \text{ J K}^{-1} \text{ mol}^{-1} = 0.08205 \text{ atm L K}^{-1} \text{ mol}^{-1}$
Speed of light	$c = 2.9979 \times 10^8 \text{ m s}^{-1}$
Planck's constant	$h = 6.6261 \times 10^{-34} \text{ J s}$
Faraday's constant	$F = 9.6485 \times 10^4 \text{ C mol}^{-1}$
Mass of electron	$m_e = 9.1093 \times 10^{-31} \text{ kg}$
Standard pressure	$P = 1 \text{ bar} = 10^5 \text{ Pa}$
Atmospheric pressure	$P_{\text{atm}} = 1.01325 \times 10^5 \text{ Pa} = 760 \text{ mmHg} = 760 \text{ torr}$
Zero of the Celsius scale	$273.15 \text{ K}$
1 picometer (pm)	$10^{-12} \text{ m}; 1 \text{ Å} = 10^{-10} \text{ m}$
1 nanometer (nm)	$10^{-9} \text{ m}$
	$1 \text{ eV} = 1.6021 \times 10^{-19} \text{ J}$
	$1 \text{ cal} = 4.184 \text{ J}$
	$1 \text{ amu} = 1.6605 \times 10^{-27} \text{ kg}$
Charge of an electron	$1.6021 \times 10^{-19} \text{ C}$
Ideal gas equation	$PV = nRT$



## Physical Constants and Equations

Enthalpy	$H = U + PV$
Gibbs free energy	$G = H - TS$
	$\Delta_r G = \Delta G^0 + RT \ln Q$
	$\Delta_r G^0 = -RT \ln K = -nFE_{cell}^0$
Entropy change	$\Delta S = \frac{q_{rev}}{T}$ , where $q_{rev}$ is heat for the reversible process
Entropy change	$\Delta S = nR \ln \frac{v_2}{v_1}$ (for isothermal expansion of an ideal gas)
Nernst equation	$E = E^0 + \frac{RT}{nF} \ln \frac{C_{oxidation}}{C_{reduction}}$
Energy of a photon	$E = \frac{hc}{\lambda}$
Integrated rate law	
Zeroth-order	$[A] = [A]_0 - kt$
First-order	$\ln [A] = \ln [A]_0 - kt$
Second order	$\frac{1}{[A]} = \frac{1}{[A]_0} + kt$
Arrhenius equation	$k = Ae^{-E_a/RT}$
Equation of linear calibration curve	$y = mx + n$
Lambert-Beer equation	$A = \varepsilon lc$



## Periodic Table of Elements

1																	18																		
<div>1 H 1.008</div>		2		<div><div>atomic number</div><div>Symbol</div><div>atomic weight</div></div>														13		14		15		16		17		<div>2 He 4.003</div>							
<div>3 Li 6.94</div>		<div>4 Be 9.01</div>																<div>5 B 10.81</div>		<div>6 C 12.01</div>		<div>7 N 14.01</div>		<div>8 O 16.00</div>		<div>9 F 19.00</div>		<div>10 Ne 20.18</div>							
<div>11 Na 22.99</div>		<div>12 Mg 24.31</div>		3		4		5		6		7		8		9		10		11		12		<div>13 Al 26.98</div>		<div>14 Si 28.09</div>		<div>15 P 30.97</div>		<div>16 S 32.06</div>		<div>17 Cl 35.45</div>		<div>18 Ar 39.95</div>	
<div>19 K 39.10</div>		<div>20 Ca 40.08</div>		<div>21 Sc 44.96</div>		<div>22 Ti 47.87</div>		<div>23 V 50.94</div>		<div>24 Cr 52.00</div>		<div>25 Mn 54.94</div>		<div>26 Fe 55.85</div>		<div>27 Co 58.93</div>		<div>28 Ni 58.69</div>		<div>29 Cu 63.55</div>		<div>30 Zn 65.38</div>		<div>31 Ga 69.72</div>		<div>32 Ge 72.63</div>		<div>33 As 74.92</div>		<div>34 Se 78.97</div>		<div>35 Br 79.90</div>		<div>36 Kr 83.80</div>	
<div>37 Rb 85.47</div>		<div>38 Sr 87.62</div>		<div>39 Y 88.91</div>		<div>40 Zr 91.22</div>		<div>41 Nb 92.91</div>		<div>42 Mo 95.95</div>		<div>43 Tc -</div>		<div>44 Ru 101.1</div>		<div>45 Rh 102.9</div>		<div>46 Pd 106.4</div>		<div>47 Ag 107.9</div>		<div>48 Cd 112.4</div>		<div>49 In 114.8</div>		<div>50 Sn 118.7</div>		<div>51 Sb 121.8</div>		<div>52 Te 127.6</div>		<div>53 I 126.9</div>		<div>54 Xe 131.3</div>	
<div>55 Cs 132.9</div>		<div>56 Ba 137.3</div>		57-71		<div>72 Hf 178.5</div>		<div>73 Ta 180.9</div>		<div>74 W 183.8</div>		<div>75 Re 186.2</div>		<div>76 Os 190.2</div>		<div>77 Ir 192.2</div>		<div>78 Pt 195.1</div>		<div>79 Au 197.0</div>		<div>80 Hg 200.6</div>		<div>81 Tl 204.4</div>		<div>82 Pb 207.2</div>		<div>83 Bi 209.0</div>		<div>84 Po -</div>		<div>85 At -</div>		<div>86 Rn -</div>	
<div>87 Fr -</div>		<div>88 Ra -</div>		89-103		<div>104 Rf -</div>		<div>105 Db -</div>		<div>106 Sg -</div>		<div>107 Bh -</div>		<div>108 Hs -</div>		<div>109 Mt -</div>		<div>110 Ds -</div>		<div>111 Rg -</div>		<div>112 Cn -</div>		<div>113 Nh -</div>		<div>114 Fl -</div>		<div>115 Mc -</div>		<div>116 Lv -</div>		<div>117 Ts -</div>		<div>118 Og -</div>	

57 La 138.9	58 Ce 140.1	59 Pr 140.9	60 Nd 144.2	61 Pm -	62 Sm 150.4	63 Eu 152.0	64 Gd 157.3	65 Tb 158.9	66 Dy 162.5	67 Ho 164.9	68 Er 167.3	69 Tm 168.9	70 Yb 173.0	71 Lu 175.0
89 Ac -	90 Th 232.0	91 Pa 231.0	92 U 238.0	93 Np -	94 Pu -	95 Am -	96 Cm -	97 Bk -	98 Cf -	99 Es -	100 Fm -	101 Md -	102 No -	103 Lr -

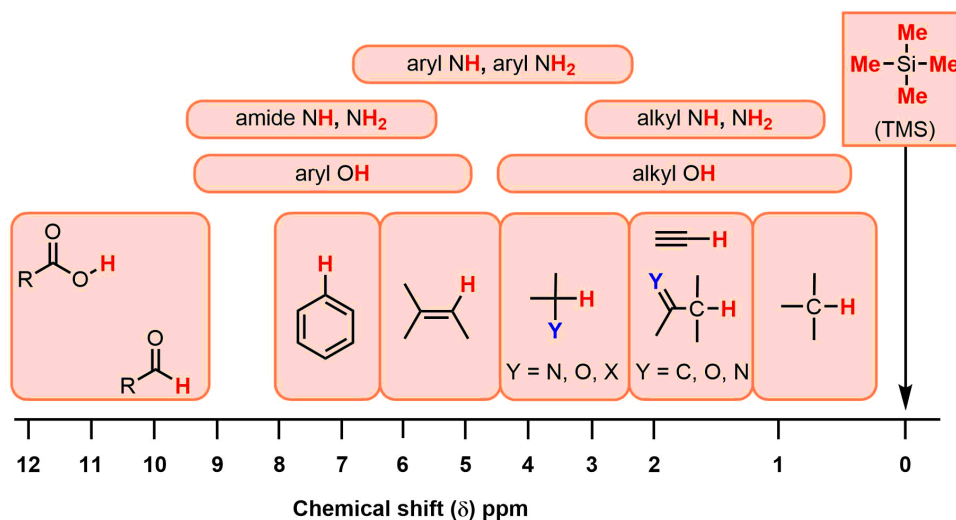


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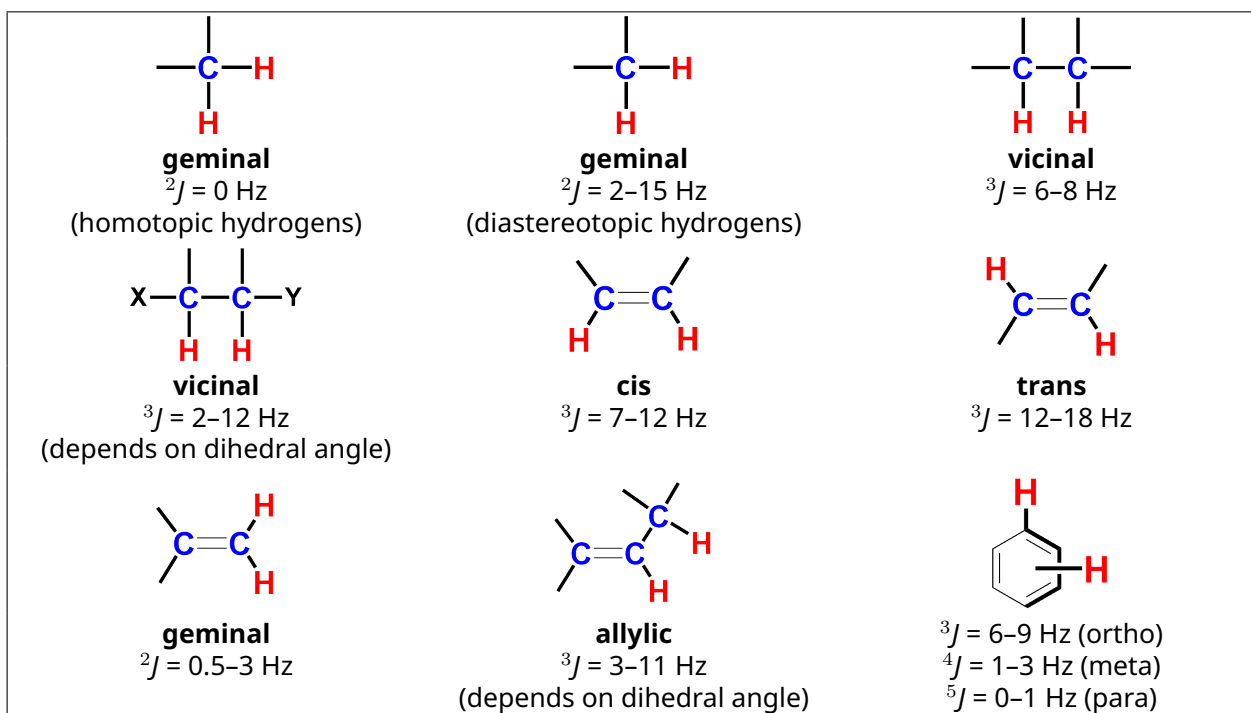
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## <sup>1</sup>H-NMR Chemical Shifts

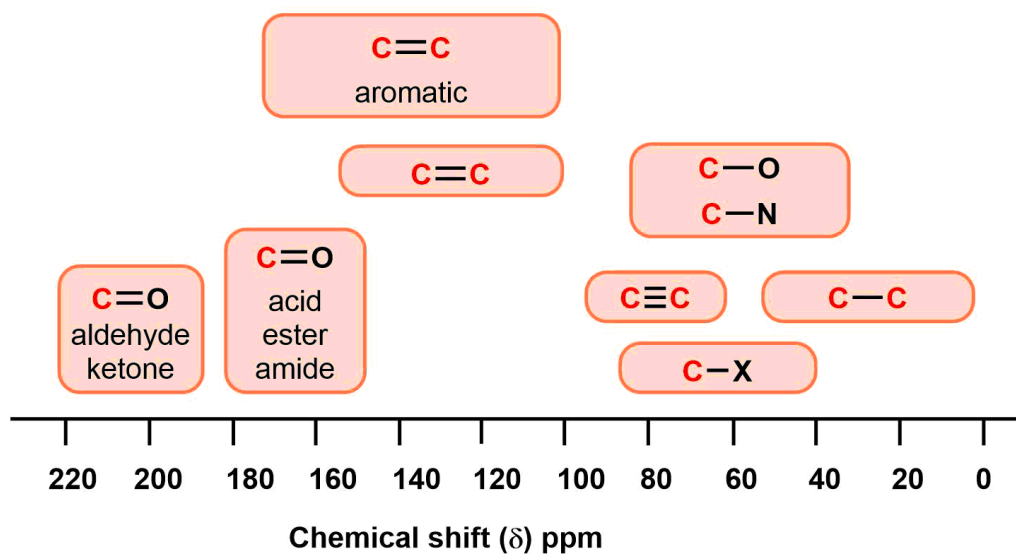


## Typical Coupling Constants





## <sup>13</sup>C-NMR Chemical Shifts







## IR Absorption Frequency Table

Functional Group	Type of Vibration	Absorption Region (cm <sup>-1</sup> )	Intensity
<b>Alcohol</b>			
O-H	(stretch, H-bonded)	3600–3200	strong, broad
	(stretch, free)	3700–3500	strong, sharp
C-O	(stretch)	1150–1050	strong
<b>Alkane</b>			
C-H	stretch	3000–2850	strong
	bending	1480–1350	variable
<b>Alkene</b>			
=C-H	stretch	3100–3010	medium
	bending	1000–675	strong
C=C	stretch	1680–1620	variable
<b>Alkyl Halide</b>			
C-F	stretch	1400–1000	strong
C-Cl	stretch	800–600	strong
C-Br	stretch	600–500	strong
C-I	stretch	500	strong
<b>Alkyne</b>			
C-H	stretch	3300	strong, sharp
C≡C	stretch	2260–2100	variable, not present in symmetrical alkynes



## IR Absorption Frequency Table

<b>Amine</b>			
N-H	stretch	3500-3300	medium (primary amines have two bands; secondary amines have one band, often very weak)
C-N	stretch	1360-1080	medium-weak
N-H	bending	1600	medium
<b>Aromatic</b>			
C-H	stretch	3100-3000	medium
C=C	stretch	1600-1400	medium-weak, multiple bands
<b>Carbonyl</b>			
C=O	stretch	1820-1670	strong
<b>Acid</b>			
C=O	stretch	1725-1700	strong
O-H	stretch	3300-2500	strong, very broad
C-O	stretch	1320-1210	strong
<b>Aldehyde</b>			
C=O	stretch	1740-1720	strong
C-H	stretch	2850-2820 & 2750-2720	medium, two peaks
<b>Amide</b>			
C=O	stretch	1690-1640	strong
N-H	stretch	3500-3100	unsubstituted have two bands
	bending	1640-1550	



## IR Absorption Frequency Table

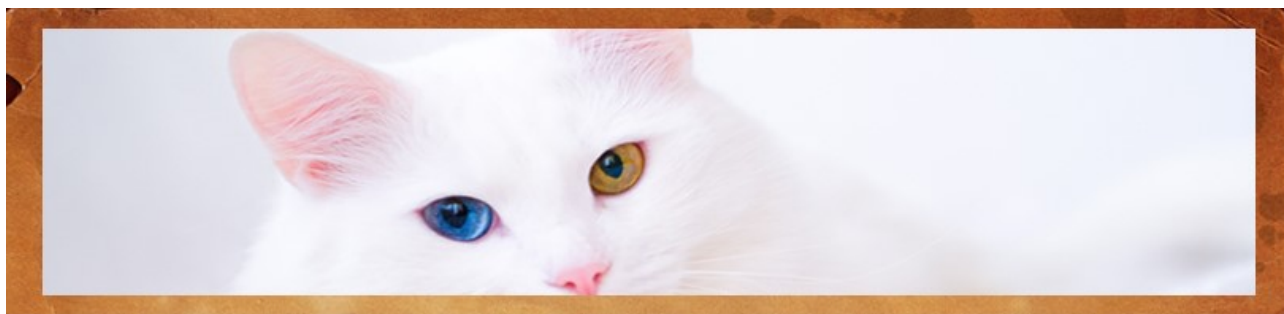
<b>Anhydride</b>			
C=O	stretch	1830-1800 & 1775-1740	two bands
<b>Ester</b>			
C=O	stretch	1750-1735	strong
C-O	stretch	1300-1000	two bands or more
<b>Ketone</b>			
acyclic	stretch	1725-1705	strong
cyclic	stretch	3-membered - 1850	strong
	stretch	4-membered - 1780	strong
	stretch	5-membered - 1745	strong
	stretch	6-membered - 1715	strong
	stretch	7-membered - 1705	strong
$\alpha, \beta$ -unsaturated	stretch	1685-1665	strong
conjugation moves absorptions to lower wavenumbers			
aryl ketone	stretch	1700-1680	strong
<b>Ether</b>			
C-O	stretch	1300-1000 (1150-1070)	strong
<b>Nitrile</b>			
C $\equiv$ N	stretch	2260-2210	medium
<b>Nitro</b>			
N-O	stretch	1560-1515 & 1385-1345	strong, two bands



Q1-1

Icelandic (Iceland)

## Two Beauties of Turkey: the Van Cat and the Ankara Cat



The most beautiful of cats, the Van cat is a pure breed living only in Lake Van basin. Another endemic cat breed is the Ankara cat. They are called Angora cats. Their most important feature is their two different eye colors.

Van cat

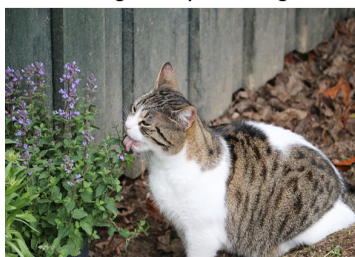


Ankara cat

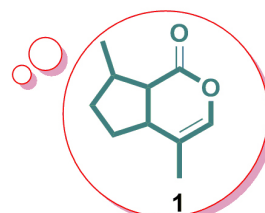
*Nepeta cataria* (catnip)

Just like people, cats can sometimes be stressed and angry. Just as people are made happy by melatonin, the stress of cats can be reduced and they can be made happy thanks to a natural product. Nepetalactone is an organic compound isolated from the plant catnip (*Nepeta cataria*), which acts as a cat attractant. Nepetalactone is a ten-carbon bicyclic monoterpene compound derived from isoprene with two fused rings: a cyclopentane and a lactone.

Cat eating catnip in the garden



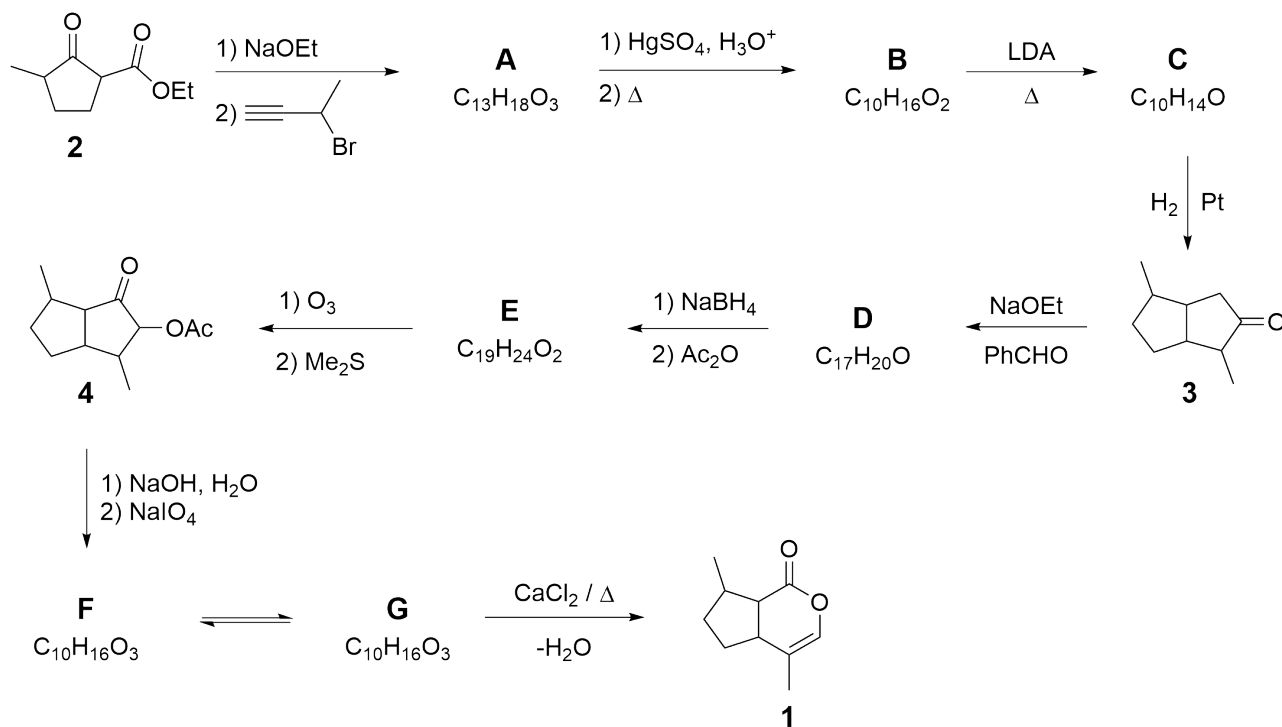
Cat's dream



Nepetalactone



## Total synthesis of nepetalactone:



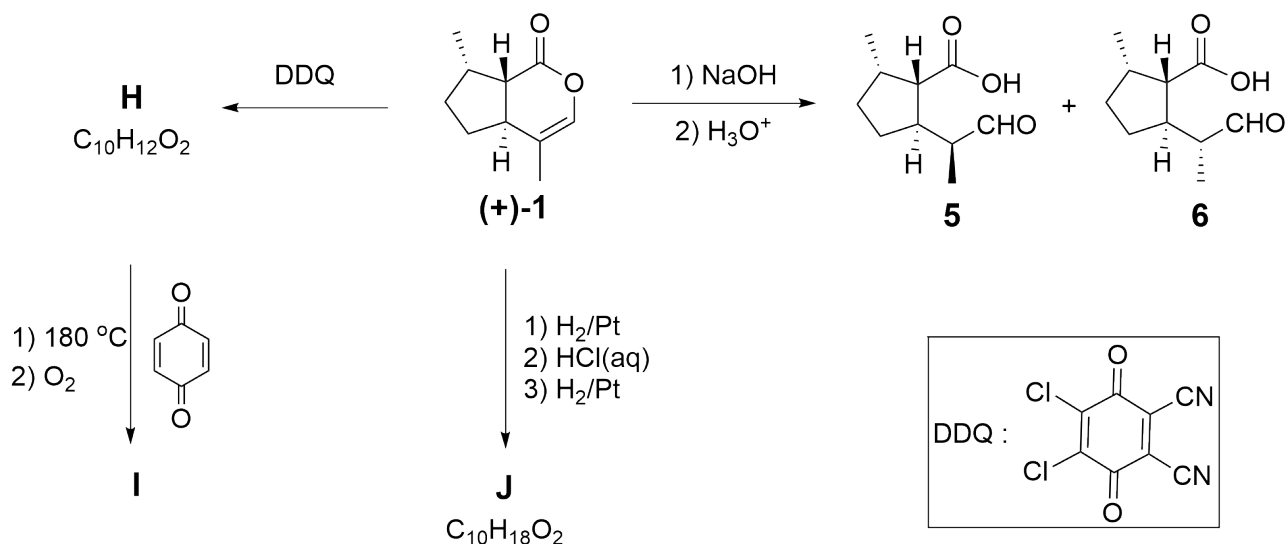
**1.1** The above scheme describes the total synthesis of nepetalactone. **Draw** structures of **A–G**, without stereochemical details. 14.0pt

**Hints:**

- Compound **A** has strong and sharp band at  $3300\text{ cm}^{-1}$  in the IR spectrum.
- **A**, **B**, and **F** are monocyclic, while **C**, **D**, **E**, and **G** are bicyclic compounds.
- **F** has one doublet at  $\sim 9.8\text{ ppm}$  in the  $^1H$ -NMR spectrum.



## Reactions of nepetalactone:



The above scheme includes a few reactions of one of the enantiopure (handhverfuhreint) nepetalactone **1** isomers. Three of the reaction products (**5**, **6**, and **J**) are used as insect repellents in industry.

- 1.2** For the relationship between **5** and **6**, which of the following is/are true? Tick 4.0pt  
the box next to the correct answer(s) on your answer sheets.

Reaction of **1** with DDQ gives highly conjugated compound **H**. Also, thermal reaction of compound **H** with *p*-quinone gives **I** with molar mass of 226.28 g/mol.

- 1.3** Draw the structures of **H**, **I**, and **J** indicating stereochemistry. 6.0pt

**Hints:**

- During the formation of **I**, sequential pericyclic reactions (hringmyndunarhvörf) and an oxidation reaction (due to the presence of  $O_2$ ) take place, and a well-known gas forms during the reaction.
- J** has a strong and very broad band between  $3300$  and  $2500\text{ cm}^{-1}$  in the IR spectrum.



## Two Beauties of Turkey: the Van Cat and the Ankara Cat

1.1 (14.0 pt)

<b>A</b>	<b>B</b>
<b>C</b>	<b>D</b>
<b>E</b>	<b>F</b>
<b>G</b>	



# A1-2

Icelandic (Iceland)

**1.2** (4.0 pt)

- ☐ Enantiomers (handhverfur)
- ☐ Diastereomers (fjölhverfur)
- ☐ Identical (eins)
- ☐ Stereoisomers (rúmhverfur)

**1.3** (6.0 pt)

<b>H</b>	<b>I</b>
<b>J</b>	

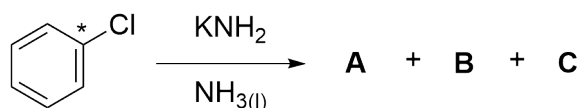




## A Tale of a Reactive Intermediate

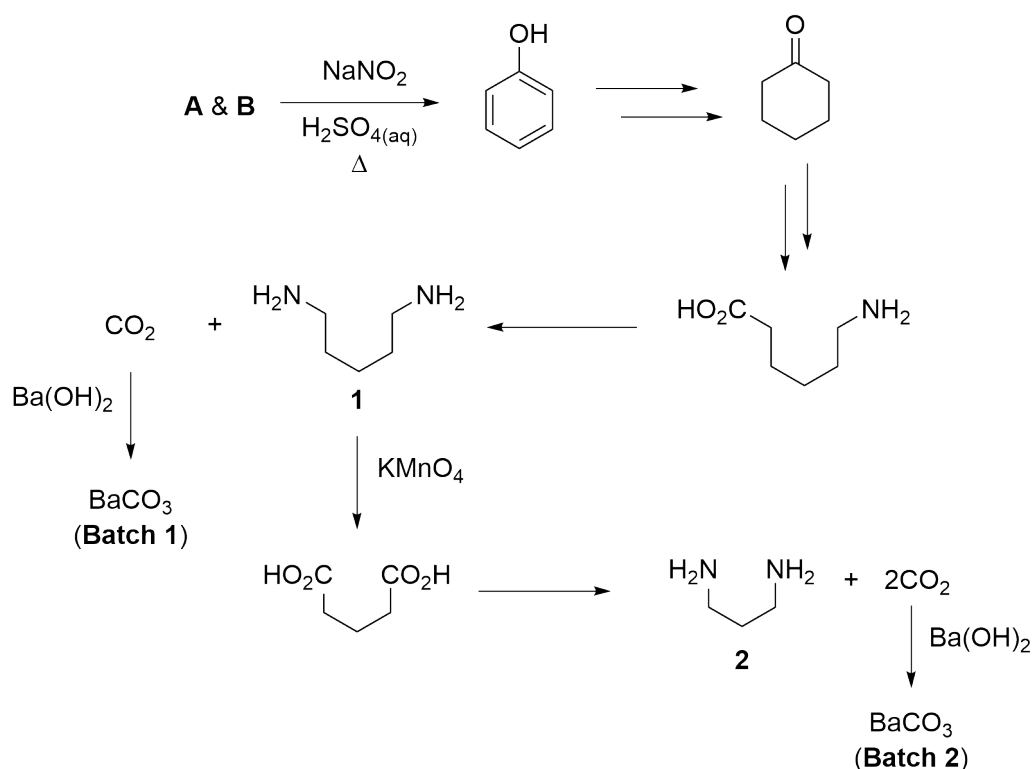
Arynes constitute a special class of reactive intermediates (hvarfgjörn milliefni). The first experimental evidence for the structure of an aryne (benzyne) was demonstrated in 1953 via the elegant labeling experiments by John D. Roberts and coworkers.

In one such experiment, chlorobenzene, whose carbon at position 1 was labeled with radioactive  $^{14}\text{C}$ , was reacted with  $\text{KNH}_2$  in liquid  $\text{NH}_3$  to give nearly equal amounts of isotopic isomers **A** and **B** along with the inorganic salt **C**. This reaction proceeds via the formation of aryne intermediate **D**.



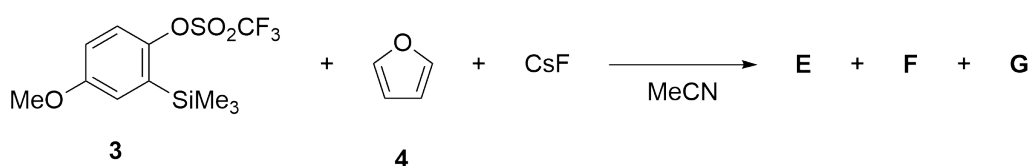
**2.1** **Draw** the structures of **A**, **B** and **D**, and **provide** the formula of **C**. **Indicate** the position(s) of  $^{14}\text{C}$ -labeled carbon(s) with an asterisk (\*) whenever applicable. 7.0pt

Analysis of the  $^{14}\text{C}$ -labeled product(s) was achieved via degradation (niðurbrots-) experiments (the  $^{14}\text{C}$ -labeled carbons are not shown on the structures). Radioactivities of the intermediates and final products were examined.



**2.2** Tick the appropriate boxes on the answer sheet for the intermediates and products that you expect to exhibit radioactivity. 9.0pt

With the aim of facilitating aryne formation, Kobayashi and co-workers developed a fluoride -induced aryne generation protocol. Using this method, benzene derivative **3** is reacted with furan (**4**) in the presence of CsF, resulting in the formation of **E**, **F**, and **G**.



- Combustion analysis of **E** revealed the following atom content: 75.8% carbon, 5.8% hydrogen, and 18.4% oxygen.
- **E** does not have a proton that is exchangeable with D<sub>2</sub>O in <sup>1</sup>H-NMR spectroscopy.
- **F** is an ionic compound.

<b>2.3</b>	<b><u>Determine</u></b> the structures of <b>E</b> , <b>F</b> , and <b>G</b> (without stereochemical details).	8.0pt
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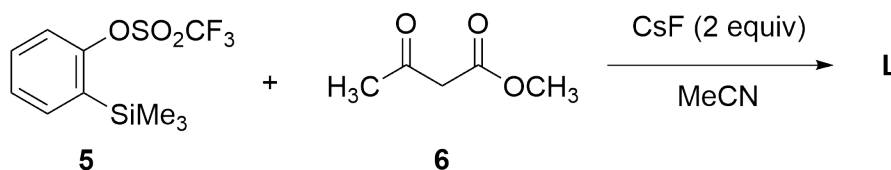
In the absence of a nucleophile or a trapping agent, arynes can undergo [2+2]-type cyclodimerization (hringtvlīdun) or [2+2+2]-type cyclotrimerization (hringþrīlīdun) reactions under suitable conditions. The aryne derivative that is obtained when **3** is treated with one equivalent of CsF in MeCN can give, in principle, four different dimerization and trimerization products (**H–K**).

- **H** has two planes of symmetry.
- **I** is expected to exhibit 21 signals in its  $^{13}\text{C}$ -NMR spectrum.
- **I** and **J** both exhibit an  $m/z$  value of 318.1 in their mass spectra.

**2.4**     **Determine** the structures of **H–K**. 16pt

When **5** is reacted with  $\beta$ -ketoester **6** in the presence of 2 equivalents (jafngildi) of CsF at 80 °C, **L** is obtained as the major product (aðalmyndefni). The  $^1\text{H}$ -NMR and  $^{13}\text{C}$ -NMR data for **L**, recorded in  $\text{CDCl}_3$ , are as follows:

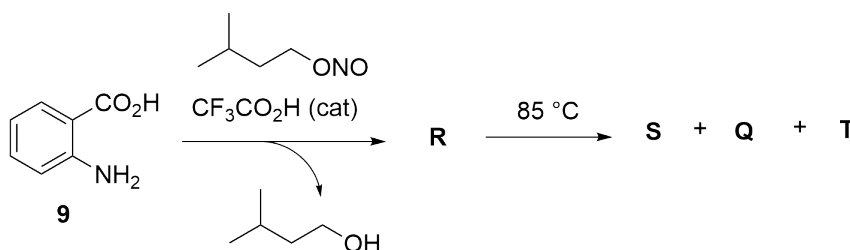
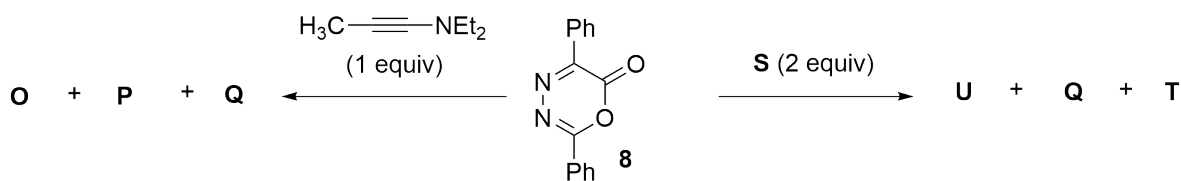
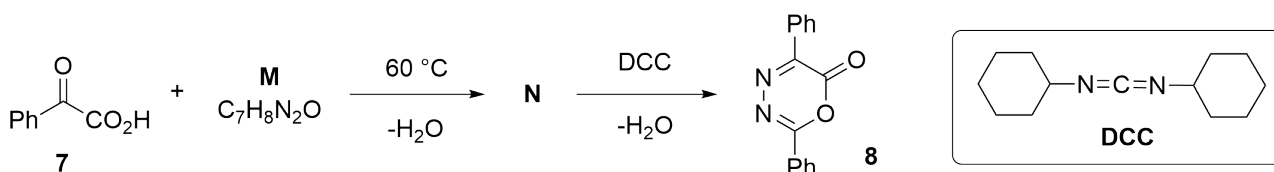
- <sup>1</sup>H-NMR: δ 7.79 (dd, *J* = 7.6, 1.5 Hz, 1H), 7.47–7.33 (m, 2H), 7.25–7.20 (m, 1H), 3.91 (s, 2H), 3.66 (s, 3H), 2.56 (s, 3H) ppm.
- <sup>13</sup>C-NMR: δ 201.3, 172.0, 137.1, 134.4, 132.8, 132.1, 130.1, 127.5, 51.9, 40.2, 28.8 ppm.



5.0pt

4.0pt

- **Q** and **T** are gases under ambient conditions.
- **O** and **P** are constitutional isomers.
- **Q** does not have any signals in its IR spectrum.
- Heating 1 mol of **R** at 85 °C generates 1 mol of reactive intermediate **S**.
- Reaction of **8** with two equivalents of **S** gives **U**, **Q**, and **T**.



cat= catalyst (hvati)

28.0pt

**A Tale of a Reactive Intermediate****2.1** (7.0 pt)

<b>A</b>	<b>B</b>
<b>C</b>	<b>D</b>

**2.2** (9.0 pt)Considering only **A**:

- ☐ Compound **1**
- ☐ BaCO<sub>3</sub> (Batch **1**)
- ☐ Compound **2**
- ☐ BaCO<sub>3</sub> (Batch **2**)

Considering only **B**:

- ☐ Compound **1**
- ☐ BaCO<sub>3</sub> (Batch **1**)
- ☐ Compound **2**
- ☐ BaCO<sub>3</sub> (Batch **2**)



**2.3** (8.0 pt)

**E**

**F**

**G**

**2.4** (16.0 pt)

**H**

**I**

**J**

**K**



**2.5** (5.0 pt)

**L**

**2.6** (4.0 pt)

- ☐  $\text{F}^-$  hydrolyzes the trifluoromethanesulfonate ( $\text{O}_3\text{SCF}_3$ ) group of **5**.
- ☐  $\text{F}^-$  attacks the  $-\text{SiMe}_3$  group of **5**.
- ☐  $\text{F}^-$  acts as a base to deprotonate of **6**.
- ☐  $\text{F}^-$  acts as a nucleophile and attacks the ester group of **6**.

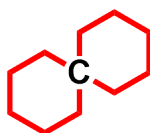


2.7 (28.0 pt)

M	N
O and P	Q
R	S
T	U

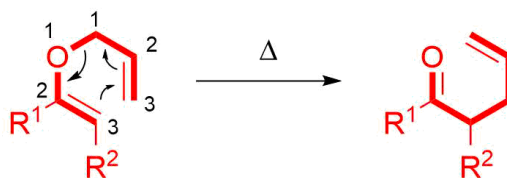
**(±)-Coerulescine**

A spiro compound is typically an organic compound containing rings linked together by one common atom (spiroatom) as carbon atom with bold in figure below. The spiro[pyrrolidin-3,3'-oxindole] ring system is a structural framework incorporated in several cytostatic alkaloids and unnatural compounds. Coerulescine (**1**) and horsfiline are the simplest prototype members of this subfamily that show diverse biological activity and can be synthesized by the route shown below.

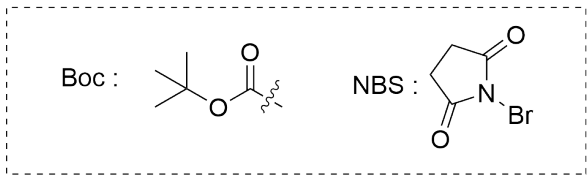


Claisen rearrangement (umröðun), which is a [3,3]-sigmatropic rearrangement, is a powerful carbon-carbon bond-forming reaction in which an allyl vinyl ether is converted thermally to an unsaturated carbonyl compound as shown in the Scheme below. When compound **A** is heated, it undergoes Claisen rearrangement to give carbonyl compound **B**.

**For this entire task, your answers can be given without any stereochemical details.**









3.1 **Draw** the structures of **A** and **B**.

8.0pt

- **A** is an inseparable mixture of *cis/trans* isomers.
- **B** has IR absorption at  $1726\text{ cm}^{-1}$ .

3.2 **Draw** structures for **C**, **D**, **E**, and **F**.

16.0pt

- **D–F** have a bicyclic structure.

3.3 **Choose** the correct order of steps for the transformation of **F** to **G**.

4.0pt

3.4 **Draw** structures for **G** and **H** (both spiro compounds).

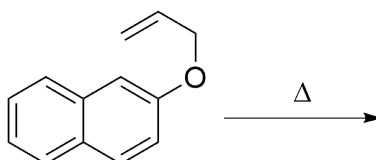
8.0pt

3.5 **Draw** the structure of the intermediate produced by treatment with *n*-BuLi in the step **H** → **coerulescine (1)**. 5.0pt

**Coerulescine (1)**, on treatment with *N*-bromosuccinimide (NBS), gives the bromo derivative, which upon heating with sodium methoxide in the presence of cuprous iodide gives **horsfiline (I)** in 60% yield.

3.6 **Choose** the correct structure for compound **I** consistent with the following selected  $^1\text{H-NMR}$  data:  $\delta$  7.05 (d,  $J = 1.4\text{ Hz}$ , 1H), 6.78 (d,  $J = 8.0\text{ Hz}$ , 1H), 6.72 (dd,  $J = 8.0, 1.4\text{ Hz}$ , 1H) ppm. 5.0pt

3.7 When the allyl ether of 2-naphthol is heated, a sigmatropic rearrangement is initiated. **Write** the structure of the major product isolated from this reaction. 5.0pt





## (±)-Coerulescine

**3.1** (8.0 pt)

**A**

**B**

**3.2** (16.0 pt)

**C**

**D**

**E**

**F**



**3.3** (4.0 pt)

- ☐ Imine formation, then reduction, then amidation
- ☐ Amidation, then imine formation, then reduction
- ☐ Reduction, then amidation, then imine formation

**3.4** (8.0 pt)

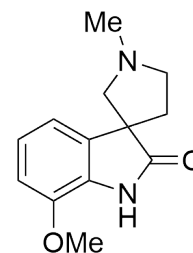
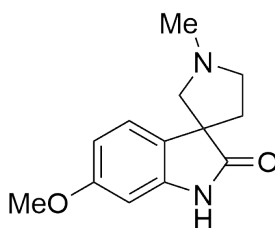
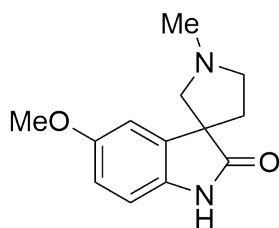
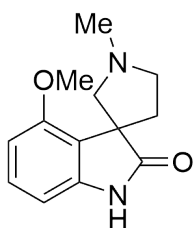
**G**

**H**

**3.5** (5.0 pt)



**3.6** (5.0 pt)

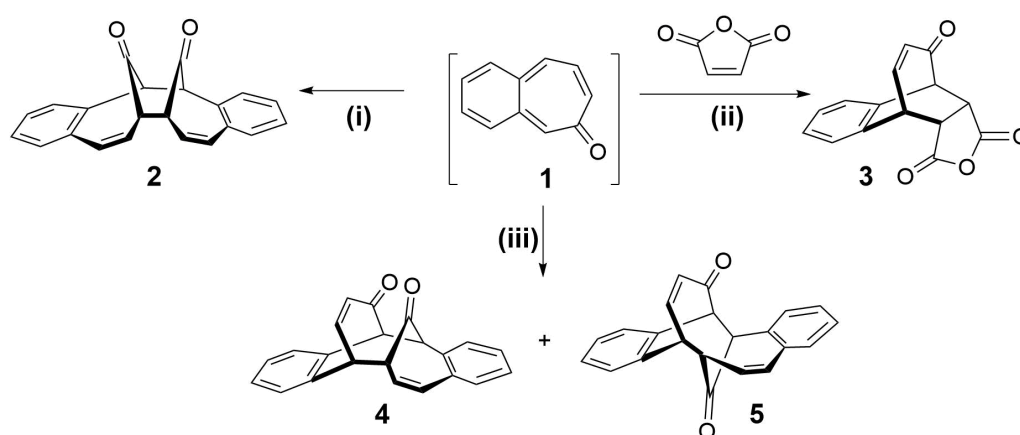


**3.7** (5.0 pt)



## Symmetry Does Matter!

There are numerous reactions in organic chemistry that proceed through cyclic transition states and these are classified as pericyclic reactions. Woodward–Hoffmann rules, developed by Robert B. Woodward and Roald Hoffmann, are used to rationalize stereochemical aspects and the activation energy of pericyclic reactions.



Woodward–Hoffmann rules				
	Electrocyclic reactions		Cycloadditions	
Number of electrons	Thermal ( $\Delta$ )	Photochemical ( $h\nu$ )	Thermal ( $\Delta$ )	Photochemical ( $h\nu$ )
$4n$ ( $n = 1, 2, \dots$ )	Conrotatory (con)	Disrotatory	Disfavored	Favored
$4n+2$ ( $n = 1, 2, \dots$ )	Disrotatory (dis)	Conrotatory	Favored	Disfavored

**4.1** Fill in the table for reactions (i)–(iii) or products 2–5:

12.0pt

There are three possible benzotropone isomers. Although two of the benzotropone isomers were isolated, 3,4-benzotropone (**1**) has not been isolated. Its instability is attributed to the *o*-quinoidal structure of **1** because it has no sextet electron system in the benzene ring.

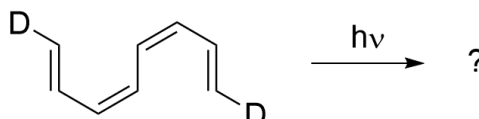
**4.2** Draw the structures of stable benzotropone isomers **A** (with 6 signals in its  $^{13}\text{C}$ -NMR) and **B** (with 11 signals in its  $^{13}\text{C}$ -NMR). 6.0pt



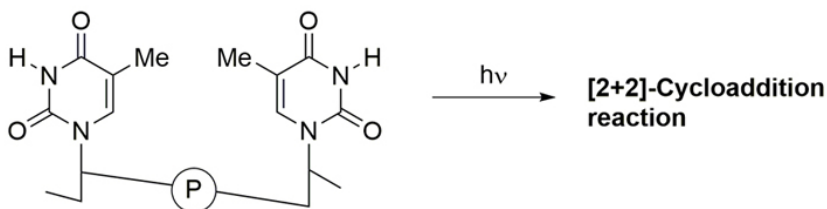
## Q4-2

Icelandic (Iceland)

- 4.3** When the following tetraene is reacted under photochemical conditions, symmetry-allowed product(s) of three different ring sizes can form according to the Woodward–Hoffmann rules. **Tick** the correct answer in each row. 6.0pt

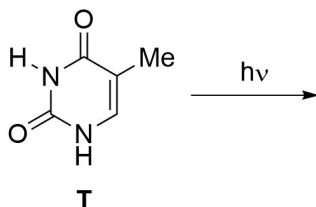


Prof. Dr. Aziz Sancar



The Nobel Prize in Chemistry 2015 was awarded jointly to the Turkish scientist Aziz Sancar, Swedish scientist Tomas Lindahl, and American scientist Paul Modrich for their “mechanistic studies of DNA repair”. Pyrimidine bases found in DNA may undergo a photochemical **[2+2]-cycloaddition reaction** (see above figure) with UV light that reaches a person’s skin, causing damage to DNA, which may ultimately lead to skin cancer. The research by Professor Aziz Sancar focused on the DNA repair mechanism for this type of damage.

Thymine (**T**) is one of the nucleobases that can undergo such a photochemical reaction with UV light. Let us assume that we have a solution of free thymine that was subjected to UV irradiation.

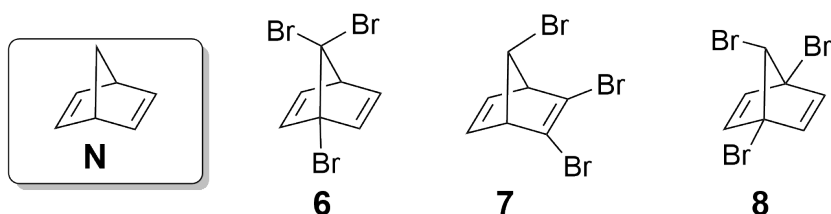


- 4.4** Considering stereochemistry, **draw** the structures of **all possible products** of this reaction between two free thymine (**T**) molecules. **Circle** the compound(s) which is/are chiral (hendin). Drawing only one enantiomer (handhverfa) of an enantiomeric pair is sufficient. Please note that only C=C bonds participate in this reaction. 16.0pt

A broad range of halogenated derivatives (afleiður) of norbornadiene (**N**) are known in the literature. Tribromo-norbornadiene (C<sub>7</sub>H<sub>5</sub>Br<sub>3</sub>) has six achiral (meso) isomers. Three of these isomers (**6**, **7**, and **8**)



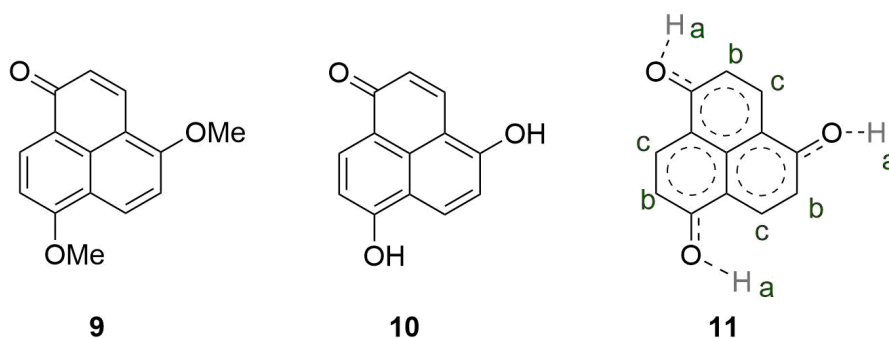
are given below.



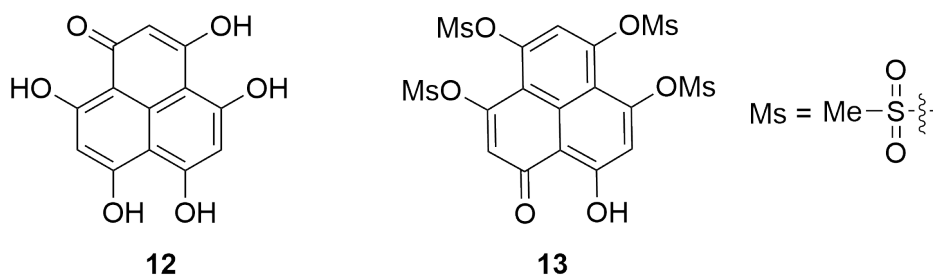
**4.5** How many signals do you expect from the  $^{13}\text{C}$ -NMR spectra of isomers **6**, **7**, and **8**? Fill in the following boxes. 9.0pt

**4.6** Draw structures of the remaining achiral (meso) tribromo-norbornadiene ( $\text{C}_7\text{H}_5\text{Br}_3$ ) achiral (meso) isomers (**C**, **D**, and **E**) in addition to **6–8** over the given figures in the boxes. 9.0pt

The NMR spectrum of ether **9** is complex. Two  $\text{MeO-}$  groups are different as are all the hydrogen atoms on the rings. However, diphenol **10** has a very simple NMR spectrum and there are only three types of protons (marked as a, b, and c). A reasonable average structure responsible for all resonance structures and its symmetry is shown as **11**.



**4.7** How many signals do you expect from the  $^{13}\text{C}$ - and  $^1\text{H}$ -NMR spectra of **12** and **13**? 8.0pt





**A4-1**

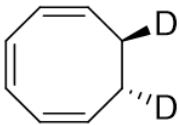
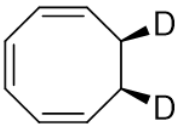
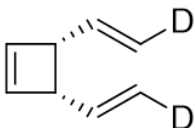
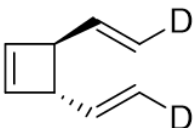
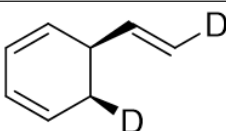
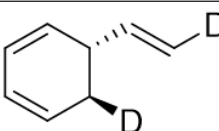
Icelandic (Iceland)

**Symmetry Does Matter!****4.1** (12.0 pt)

Reaction	Product	[? + ?] cycloaddition	$\Delta$ or $h\nu$
<b>i</b>	<b>2</b>		
<b>ii</b>	<b>3</b>		
<b>iii</b>	<b>4</b>		
	<b>5</b>		

**4.2** (6.0 pt)**A****B**

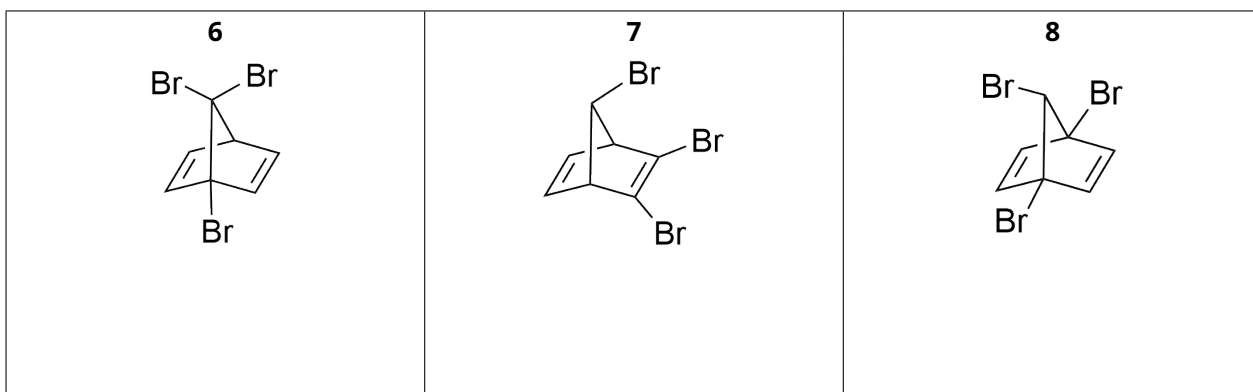
**4.3** (6.0 pt)

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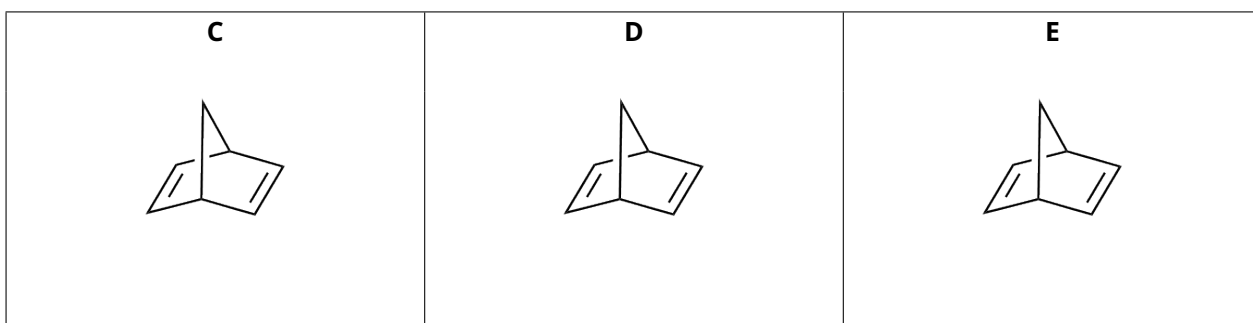
**4.4** (16.0 pt)



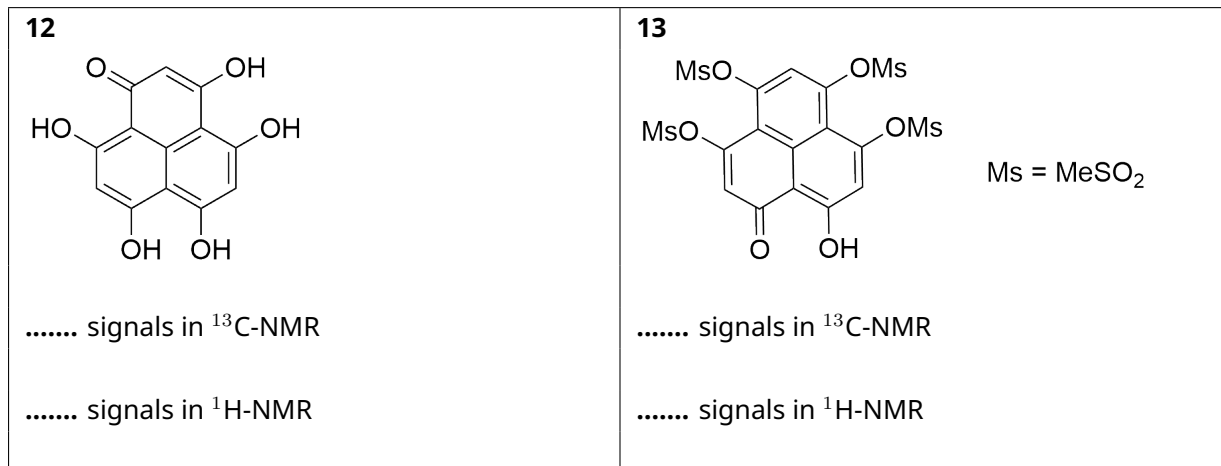
4.5 (9.0 pt)



4.6 (9.0 pt)



4.7 (8.0 pt)

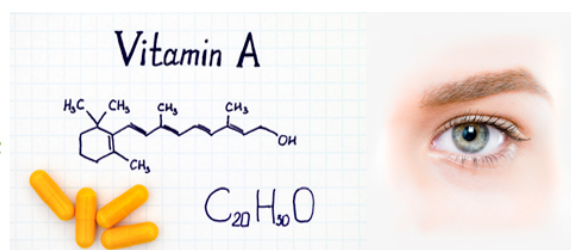




# Q5-1

Icelandic (Iceland)

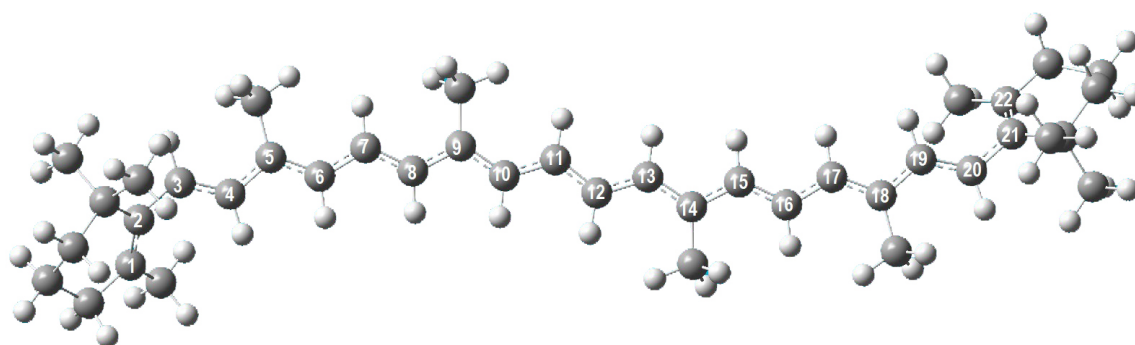
## Konya, Carrot, Beta-Carotene, Vitamin-A, Immune System, Vision



Mevlana (Rumi) was a great mystic and Sufi poet who lived out his days in Konya in the 13<sup>th</sup> century. The indirect relevance of Konya to chemistry is that the city provides 65% of the country's carrot production, from which one of the essential vitamins (vitamin A) is obtained.

Carrot is an important source of  $\beta$ -carotene, which gives the vegetable its orange color. This molecule is a red-orange pigment naturally found in plants and fruits and is a provitamin A carotenoid. It is converted to vitamin A, which is essential for normal growth and development, the immune system, and vision function.

$\beta$ -Carotene has an extended polyene chain of 22 carbon atoms. It is a conjugated  $\pi$ -system, having alternating single and double bonds. Its experimental maximum absorption wavelength ( $\lambda_{max}$ ) is 455 nm. We assume that all the bonds between  $C_1$  and  $C_{22}$  are conjugated bonds. There are 22  $\pi$ -electrons in the molecule (Figure 1).



**Figure 1.** Ball and stick representation of the structure of  $\beta$ -carotene. The gray and white spheres represent the carbon and hydrogen atoms, respectively. The numbered carbon atoms belong to the linear conjugated  $\pi$ -segment of the molecule.

To a crude approximation, the electrons in the  $C-2p_z$  orbitals, which are perpendicular to the molecular plane, are assumed to move along the entire molecule, without interacting with each other. They are like independent particles confined in a molecule moving along the x-axis in one dimension. These characteristics of  $\pi$ -electrons make them eligible for being treated by the simplest model called the **particle in one-dimensional box** model.



# Q5-2

Icelandic (Iceland)

The wavefunction and the energies of the quantized levels for an electron moving in a one-dimensional box with infinite potential walls are given as follows:

$$\Psi_n(x) = \sqrt{\frac{2}{L}} \sin \frac{n\pi x}{L} \quad (\text{Eq.1})$$

where  $n$  is the quantum number,  $n=1,2,3,4,\dots, \infty$ , and  $L$  is the box length.

$$E_n = \frac{n^2 h^2}{8m_e L^2} \quad (\text{Eq.2})$$

In two dimensions, within the framework of independent particle approximation, the wavefunction is expressed as a product of one-dimensional wavefunctions, and the energy is expressed as a sum of one-dimensional energies. The energy levels of the two dimensional rectangular box is given as follows:

$$E_{n_x, n_y} = \left[ \frac{n_x^2}{L_x^2} + \frac{n_y^2}{L_y^2} \right] \left\{ \frac{h^2}{8m_e} \right\} \quad (\text{Eq.3})$$

where  $n_x, n_y$  are the quantum numbers and they are positive integers.  $L_x, L_y$  are the dimensions of the box in the 2D model. They are positive numbers.

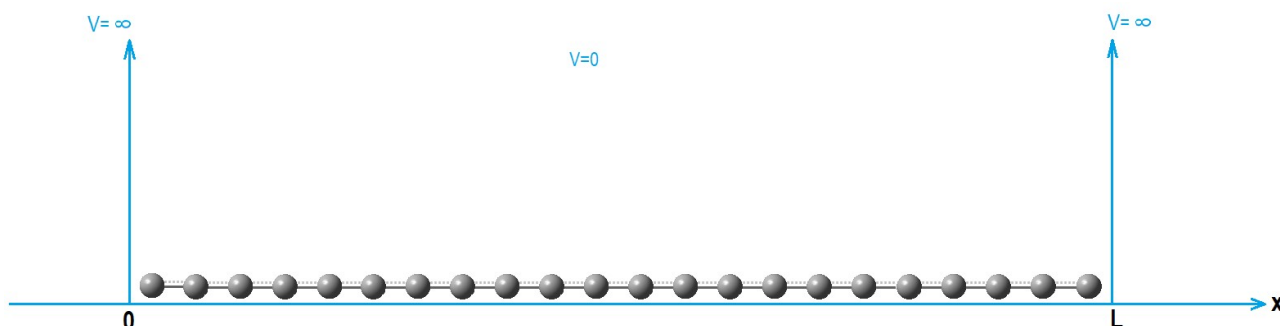
**5.1** Which two of the sentences given below are correct? **Tick** only one answer. 13.0pt

**The  $\beta$ -carotene molecule is orange in color because:**

- i) it absorbs in the visible region of the electromagnetic spectrum.
- ii) HOMO  $\rightarrow$  LUMO transition occurs by absorption of IR photon.
- iii) the spacing between the 22<sup>nd</sup> and the 23<sup>rd</sup> energy levels is equal to the energy of the IR photon at the orange wavelength.
- iv) it absorbs green/blue light and it transmits red/yellow color.
- v) it absorbs in the UV-Vis region since the molecule has no net dipole moment.

Although it is highly unrealistic, let us assume that the conjugated segment of the molecule is linear and treated with the particle in a one-dimensional box model as shown in figure 2. In this case, the length of the box can be approximated as  $L=1.40 \times n_C$  (in Å), where  $n_C$  is the number of carbon atoms in the conjugated segment.

**Use this information to answer the questions 5.2-5.6.**



**Figure 2.** Schematic representation of the conjugated line segment made up of carbon atoms of  $\beta$ -carotene in a one-dimensional box of length  $L$ .



**5.2** **Calculate** the energies (in J) of the lowest two energy levels 13.0pt

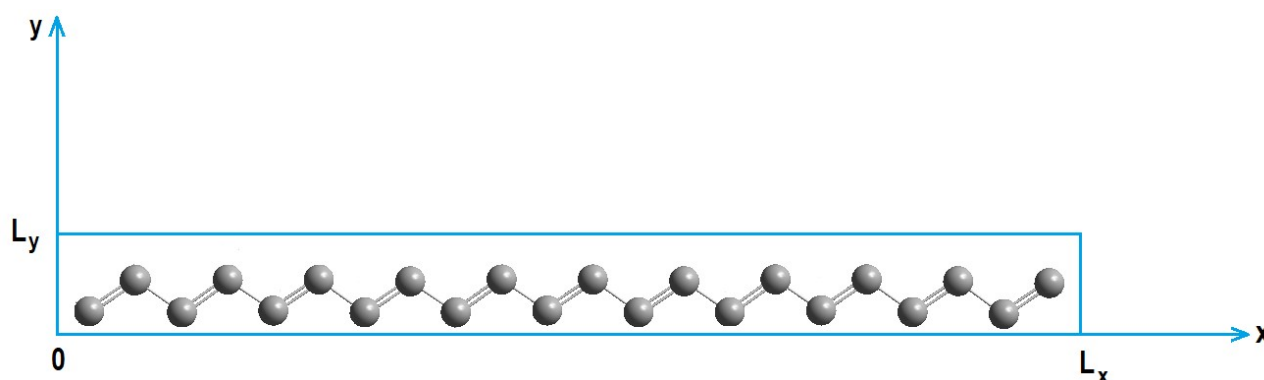
**5.3** **Draw** the wavefunctions of the lowest two energy levels with proper labelling the x-axis. 15.0pt

**5.4** **Sketch** the energy level diagram up to  $n = 4$  showing the relative spacing. 8.0pt

**5.5** What is the total  $\pi$ -energy (in J) of the molecule? 12.0pt

**5.6** **Calculate** the wavelength (in nm) at which the transition between the highest occupied and lowest unoccupied energy levels occurs. 10.0pt

Use the particle in a two-dimensional box model to answer questions 5.7–5.8.



**Figure 3.** Schematic representation of the conjugated carbon atoms of  $\beta$ -carotene in a two-dimensional box.

Assume that the conjugated segment is made up of carbon atoms that are all-*trans* to each other. The motion of the  $\pi$ -electrons is studied in the two-dimensional rectangular box with the dimensions  $L_x = 26.0 \text{ \AA}$ ,  $L_y = 3.0 \text{ \AA}$  (Figure 3).

**5.7** **Calculate** the energies (in J) of the highest occupied and the lowest unoccupied energy levels and the wavelength (in nm) at which the transition between these energy levels occurs. 17.0pt

**5.8** What should be the  $L_x$  value (in  $\text{\AA}$ ) in order for the molecule to absorb light at  $\lambda_{max} = 455 \text{ nm}$  if  $L_y$  is kept constant at  $3.0 \text{ \AA}$ . (Assume that the quantum numbers for HOMO and LUMO are the same as in the question 5.7.) 12.0pt



## Konya, Carrot, Beta-Carotene, Vitamin-A, Immune System, Vision

**5.1** (13.0 pt)

- |   |  |   |   |
|---|--|---|---|
| <input type="checkbox"/> <b>a)</b> i and ii   | <input type="checkbox"/> <b>b)</b> i and iii | <input type="checkbox"/> <b>c)</b> i and iv | <input type="checkbox"/> <b>d)</b> i and v    |
| <input type="checkbox"/> <b>e)</b> ii and iii | <input type="checkbox"/> <b>f)</b> ii and iv | <input type="checkbox"/> <b>g)</b> ii and v | <input type="checkbox"/> <b>h)</b> iii and iv |
| <input type="checkbox"/> <b>j)</b> iii and v  | <input type="checkbox"/> <b>k)</b> iv and v  |   |   |

**5.2** (13.0 pt)

Calculation:

**5.3** (15.0 pt)



**5.4** (8.0 pt)

**5.5** (12.0 pt)





**5.6** (10.0 pt)

calculation:

**5.7** (17.0 pt)

calculation:



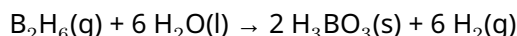
**5.8** (12.0 pt)



## Thermodynamics through an Interstellar Journey

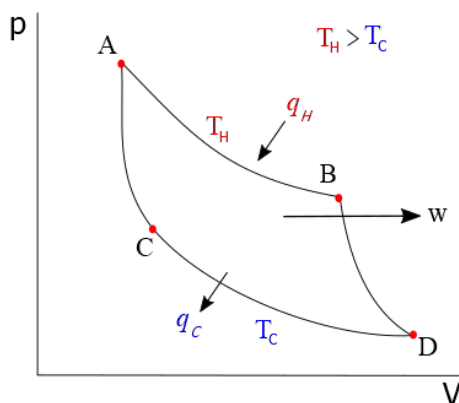
### Part 1

In a hypothetical universe, an unknown amount of diborane participates in the following reaction:



Assume that in this hypothetical universe,  $\text{H}_3\text{BO}_3(\text{s})$  obtained from this reaction was completely sublimed (þurrugað) at 300 K. The necessary energy for sublimation (þurrgun) was obtained through work released by **one cycle** of an ideal heat engine in which one mole of monoatomic perfect gas flows through the cycle described in the pressure (p) – volume (V) diagram below:

- A → B; isothermal (jafnhita) reversible expansion receiving 250 J by heat transfer ( $q_H$ ) at a temperature of 1000 K ( $T_H$ ) from a hot source.
- B → D; reversible adiabatic (óvermin) expansion.
- D → C; isothermal reversible compression (þjöppun) at a temperature of 300 K ( $T_C$ ) releasing some amount of heat ( $q_C$ ) to a cold sink.
- C → A; reversible adiabatic compression.

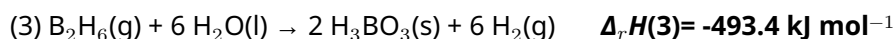
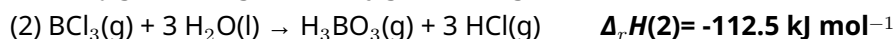
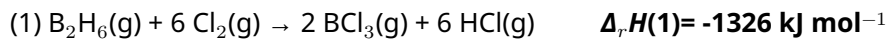


After heat transfers, the remaining energy is released as work (w). Also,  $q_H$  and  $q_C$  are related to  $T_C$  and  $T_H$  as follows:

$$\frac{|q_H|}{|q_C|} = \frac{T_H}{T_C}$$

The efficiency (nýtni) of the cycle can be found by work released by cycle (w) divided by heat absorbed by cycle ( $q_H$ ).

You are provided with the change in enthalpies of the following reactions at 300 K.



**6.1** **Calculate** the molar enthalpy of sublimation (in  $\text{kJ mol}^{-1}$ ) for  $\text{H}_3\text{BO}_3$  at 300 K. 5.0pt

**6.2** **Calculate** the  $\Delta_r U$  (internal energy) in terms (einingunum) of  $\text{kJ mol}^{-1}$  at 300 K for the reactions (2) and (4) given above (assume ideal gas behavior for each gaseous species in each reaction). 12.0pt

**6.3** **Calculate** the amount of overall work produced by a heat engine ( $|w|$ ) in terms of J and the amount of overall heat released to the cold sink ( $|q_C|$ ) in terms of J. 6.0pt

**6.4** **Calculate** the efficiency of the heat engine described above. 3.0pt

**6.5** **Calculate** the entropy (óreiðu) change ( $\Delta S$ ) for  $A \rightarrow B$  and  $D \rightarrow C$  processes in the heat engine in terms of  $\text{J K}^{-1}$ . 6.0pt

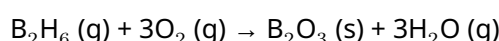
**6.6** **Calculate** the Gibbs energy change ( $\Delta G$ ) in terms of J for  $A \rightarrow B$  and  $D \rightarrow C$  processes in the heat engine. 6.0pt

**6.7** **Calculate** the ratio of pressure at point A to the pressure at point B in the cycle (standard pressure: 1 bar). 5.0pt

**6.8** **Calculate** the amount of  $\text{H}_2(\text{g})$  (in moles) produced according to the reaction given at the beginning of the task for one cycle of the engine. 3.0pt

## Part 2

Interstellar journeys can be done by using diborane as rocket fuel. Combustion of diborane is shown below:



Combustion of diborane is experimented (framkvæmd) in a 100 L closed container at different temperatures and the equilibrium conditions were recorded.



# Q6-3

Icelandic (Iceland)

	8930 K	9005 K
$B_2H_6(g)$	0.38 mol	0.49 mol
$H_2O$	0.20 mol	0.20 mol

Partial pressure of  $O_2(g)$  was stabilized to 1 bar and kept constant at all conditions. Assume that in this hypothetical universe;  $\Delta_r S^\circ$  and  $\Delta_r H^\circ$  are independent of temperature, the standard molar entropy ( $S^\circ$ ) of  $B_2O_3(s)$  does not change with pressure, all the gas species behave as an ideal gas, and all species remain in the same phase, without any further decomposition before or after reaction, at all temperatures then:

**6.9 Calculate**  $K_p$  (pressure based equilibrium constant) at 8930 K and 9005. 8.0pt

**6.10 Calculate**  $\Delta_r G^\circ$  of the reaction in terms of  $kJ\ mol^{-1}$  at 8930 K and 9005 K. (If you failed to find  $K_p$ , please use  $K_p(8930\ K) = 2$ ,  $K_p(9005\ K) = 0.5$ ) 6.0pt

**6.11 Calculate**  $\Delta_r G^\circ$  (in terms of  $kJ\ mol^{-1}$ ),  $\Delta_r H^\circ$  (in terms of  $kJ\ mol^{-1}$ ), and  $\Delta_r S^\circ$  (in terms of  $J\ mol^{-1}\ K^{-1}$ ) of the combustion reaction at 298 K. (If you failed to find  $K_p$ , please use  $K_p(8930\ K) = 2$ ,  $K_p(9005\ K) = 0.5$ ) 6.0pt

**6.12 Tick** the correct answer in the table by determining whether combustion reactions are favored or not at given T below under standard pressure (1 bar). 8.0pt

**6.13 Calculate** the  $\Delta_f H$  ( $kJ\ mol^{-1}$ ) and  $S^\circ$  ( $kJ\ mol^{-1}\ K^{-1}$ ) of  $H_2O(g)$  using the values given in the table below. ( $\Delta_f H$  = enthalpy of formation,  $S^\circ$  = standard entropy) (If you fail to find  $\Delta_r H^\circ$  and  $\Delta_r S^\circ$  of the combustion, please use  $\Delta H^\circ = 1000\ kJ\ mol^{-1}$ ,  $\Delta S^\circ = 150\ J\ K^{-1}\ mol^{-1}$ ) 6.0pt

	$\Delta_f H$ (298 K)	$S^\circ$ (298 K)
$B_2H_6(g)$	$36.40\ kJ\ mol^{-1}$	$0.23\ kJ\ mol^{-1}\ K^{-1}$
$O_2(g)$	$0.00\ kJ\ mol^{-1}$	$0.16\ kJ\ mol^{-1}\ K^{-1}$
$B_2O_3(s)$	$-1273\ kJ\ mol^{-1}$	$0.05\ kJ\ mol^{-1}\ K^{-1}$



## Thermodynamics through an Interstellar Journey

**6.1** (5.0 pt)

Show your calculation:

**6.2** (12.0 pt)

Show your calculation:

**6.3** (6.0 pt)

Show your calculation:



**6.4** (3.0 pt)

Show your calculation:

**6.5** (6.0 pt)

Show your calculation:

**6.6** (6.0 pt)

Show your calculation:



**6.7** (5.0 pt)

Show your calculation:

**6.8** (3.0 pt)

Show your calculation:





**6.9** (8.0 pt)

Show your calculation:



**6.10** (6.0 pt)

Show your calculation:

**6.11** (6.0 pt)

Show your calculation:



**6.12** (8.0 pt)

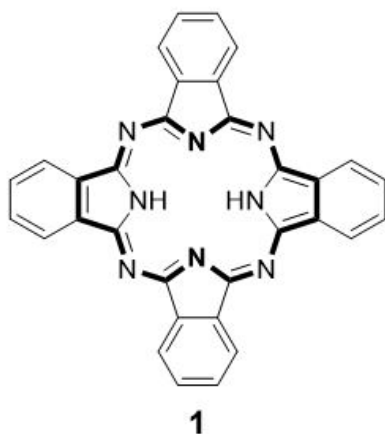
	favored	Unfavored
298 K	<input type="checkbox"/>	<input type="checkbox"/>
8930 K	<input type="checkbox"/>	<input type="checkbox"/>
9005 K	<input type="checkbox"/>	<input type="checkbox"/>
9100 K	<input type="checkbox"/>	<input type="checkbox"/>

**6.13** (6.0 pt)

Show your calculation:



## Phthalocyanines

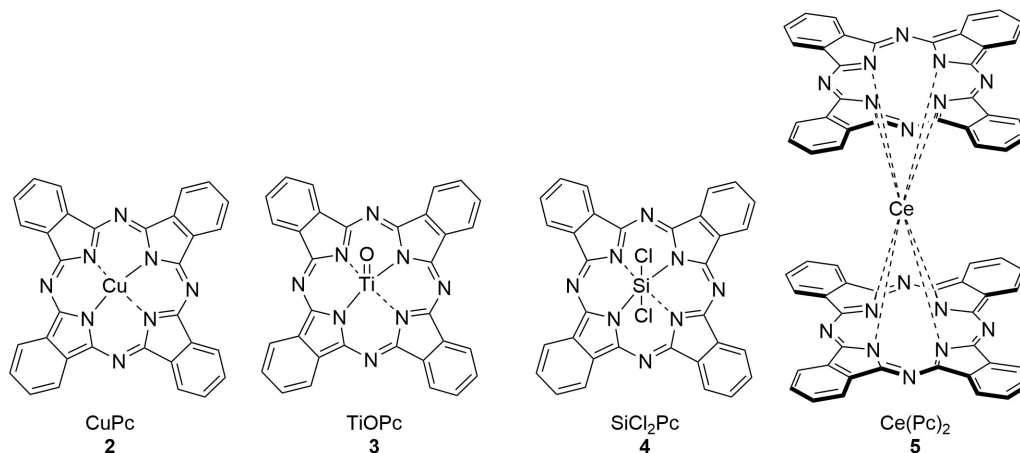


*Emeritus Professor Özer  
Bekaroğlu*

The term phthalocyanine (Pc) takes its origin from the Greek “naphtha”, which means rock oil, and “cyanine”, which means dark blue. Turkish scientist Özer Bekaroğlu can be regarded as the pioneer of Pc chemistry in Turkey.

Metal-free phthalocyanine (**1**,  $\text{H}_2\text{Pc}$ ) is a large planar macrocyclic compound with the formula  $(\text{C}_8\text{H}_4\text{N}_2)_4\text{H}_2$ .

- |            |  |       |
|------------|--|-------|
| <b>7.1</b> | How many $\pi$ -electrons are there in the bold region of the $\text{H}_2\text{Pc}$ molecule in compound <b>1</b> shown above? | 4.0pt |
|------------|--|-------|



Pcs containing one or two metal ions are called metallo-phthalocyanines (MPcs) and they exhibit different geometries as given above.

**7.2** **Complete** the table in your answer sheet by determining the coordination number of central ions in **2-5**. 8.0pt

**7.3** **Complete** the table in your answer sheet by determining the oxidation number of each metal (Cu, Ti, and Ce) in **2, 3**, and **5**. 6.0pt

**7.4** **Complete** the table in your answer sheet by determining the geometry of compounds **2-5**. 8.0pt

**7.5** **Complete** the table in your answer sheet by determining the magnetic property of compounds **2-5**. 8.0pt

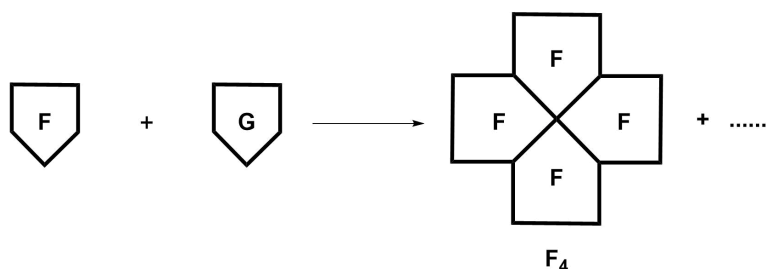
- Use the letter "**p**" for paramagnetic property and letter "**d**" for a diamagnetic property.

**7.6** **Write** the ground-state electron configuration of the silicon (Si) ion in the compound **4**, and **find** all the quantum numbers for the 2p electrons in its ground state. 14.0pt

Metal-free phthalocyanine (**1**,  $\text{H}_2\text{Pc}$ ) is commonly formed through the cyclotetramerization of phthalonitriles. On the other hand, Pcs having different substituents are called asymmetric, which can be prepared by the statistical cyclization of two different phthalonitriles. This method has no selectivity and the product is a mixture of all possible isomers.



- 7.7** **Draw** the possible products that may occur in the statistical cyclization method by using **F** and **G**. If there are any stereoisomers label as *cis*- or *trans*-. 19.0pt
- **F** and **G** represent two different symmetrical phthalonitriles.
  - One of the products is **F<sub>4</sub>** as given below.
  - Draw other products similar to the format in **F<sub>4</sub>**.



Pcs are used as photosensitizers in the photodynamic therapy (PDT) of cancer due to their strong absorption in the visible spectrum and high molar absorption coefficients. PDT consists of three essential components: **photosensitizer**, light, and oxygen. None of these is individually toxic, but together they initiate a photochemical reaction resulting in the generation of cytotoxic singlet oxygen ( $^1\text{O}_2$ ) that can destroy cancer cells.

(multiplicity)  $^1\text{O}_2$

- The multiplicity of an energy level is defined as  $2S+1$
- If the two spins are parallel ( $\uparrow\uparrow$ ),  $S = 1$ , and if the two spins are antiparallel ( $\uparrow\downarrow$ ),  $S = 0$ .

- 7.8** **Draw** the molecule orbital (MO) diagram of the lowest energy singlet state of dioxygen ( $^1\text{O}_2$ ) and calculate bond order. 12.0pt
- There are no unpaired electrons in that state!

- 7.9** If the wavelength of the light needed to excite triplet oxygen into singlet oxygen is 1270 nm, **calculate** the energy (in kJ per mole) needed for this transition process. 6.0pt



## Phthalocyanines

**7.1** (4.0 pt)

The number of  $\pi$ -electrons in an  $\text{H}_2\text{Pc}$ :

**7.2** (8.0 pt)

Central ion	Copper ion	Titanium ion	Silicon ion	Cerium ion
Coordination number				

**7.3** (6.0 pt)

Metal in compounds	2	3	5
Oxidation number			

**7.4** (8.0 pt)

Geometry	Compound
Octahedral (áttflötungur)	
Square prism (ferningsstrendingur)	
Square pyramidal (ferningspíramíti)	
Square planar (ferningslaga)	

**7.5** (8.0 pt)

Compound	Magnetic property
2	
3	
4	
5	

**7.6** (14.0 pt)

[illegible]





**7.7** (19.0 pt)

Products:



**A7-4**  
Icelandic (Iceland)

**7.8** (12.0 pt)

MO diagram:

Bond order:

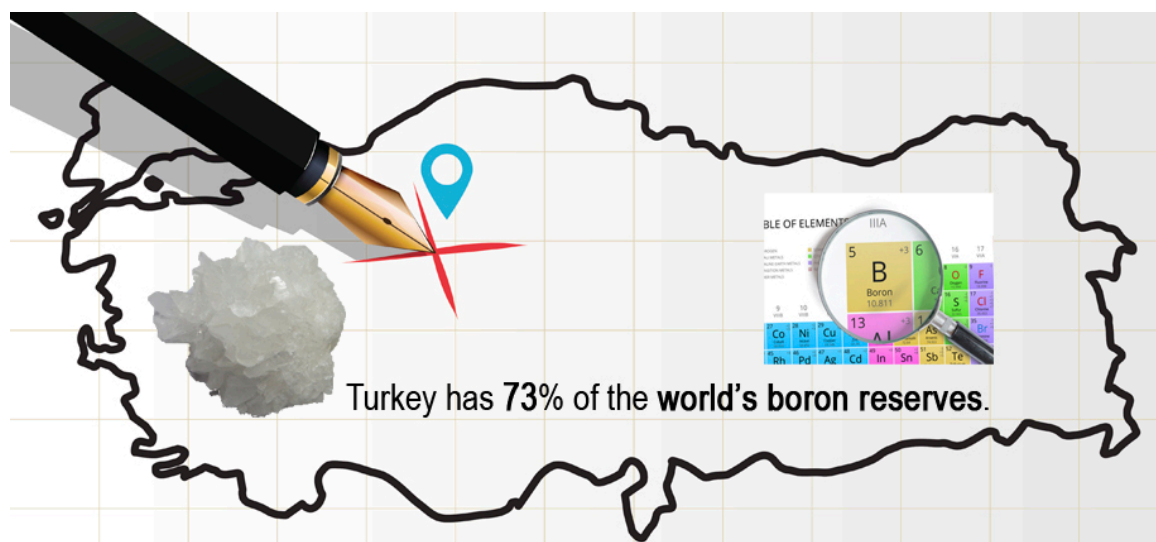
**7.9** (6.0 pt)

Show your calculation:

Energy = ..... kJ/mol



## Boron Compounds and Hydrogen Storage



Sodium borohydride ( $\text{NaBH}_4$ ) and ammonia borane ( $\text{BNH}_6$ ) are the most studied chemical hydrogen storage materials. In this question, you will explore the chemistry of boron and the use of boron compounds as hydrogen storage materials.

Borax ( $\text{Na}_2\text{B}_4\text{O}_7 \cdot n\text{H}_2\text{O}$ ) is a boron mineral that is produced by ETI Mining Company in Turkey.  $\text{NaBH}_4$  can be synthesized by the reduction of anhydrous (vatnsfríu) borax with metallic sodium under high-pressure hydrogen gas in the presence of silicon dioxide (silica) at  $700^\circ\text{C}$ , namely the Bayer process. In this process, all hydrogen is stored in  $\text{NaBH}_4$ . On the other hand, it has been shown that ammonia borane ( $\text{BNH}_6$ ) can be synthesized by the reaction of  $\text{NaBH}_4$  and ammonium sulfate in dry tetrahydrofuran (THF) at  $40^\circ\text{C}$  (**Hint:**  $\text{BNH}_6$  synthesis must be conducted in a well-ventilated fume-hood because flammable gas is generated as one of the by-products). While  $\text{NaBH}_4$  is an ionic compound, ammonia borane is a Lewis acid-base adduct.

**8.1** Write a balanced chemical equation for the synthesis of  $\text{NaBH}_4$  from anhydrous borax. 3.0pt

**8.2** Write a balanced chemical equation for the synthesis of ammonia borane from  $\text{NaBH}_4$ . 3.0pt

**8.3** Draw the molecular geometries of the  $\text{BH}_4^-$  ion and  $\text{BNH}_6$  molecule. 4.0pt

**8.4** Calculate the hydrogen content of  $\text{NaBH}_4$  and  $\text{BNH}_6$  as a percentage by mass (wt%). 4.0pt

The hydrogen stored in both compounds can be released via hydrolysis reactions in the presence of a suitable catalyst at room temperature. Upon the hydrolysis reactions, 4 and 3 moles of  $\text{H}_2$  gas are



released from the hydrolysis of 1 mole of  $\text{NaBH}_4$  and  $\text{BNH}_6$ , respectively, along with metaborate anion including B-O bonds.

**8.5** **Write** the balanced chemical equations for the hydrolysis of  $\text{NaBH}_4$  and  $\text{BNH}_6$ . 4.0pt

One of the simplest stable borates is diboron trioxide ( $\text{B}_2\text{O}_3$ ). Higher borates such as  $\text{B}_3\text{O}_6^{3-}$  having cyclic structures containing B-O-bonds can be formed. Since  $\text{B}_2\text{O}_3$  is an acidic compound, it is easily reacted with water to produce boric acid ( $\text{H}_3\text{BO}_3$ ). On the other hand, the high temperature and high-pressure reaction of  $\text{B}_2\text{O}_3$  with ammonia yields two-dimensional boron nitride, which consists of planar graphite-like sheets of alternating B and N atoms.

**8.6** **Write** the balanced chemical equations for the synthesis of boric acid and boron nitride. 4.0pt

**8.7** **Draw** the molecular structures of the  $\text{B}_3\text{O}_6^{3-}$  ion, boric acid, and a single two-dimensional boron nitride sheet. **Hint:** show at least 10 B atoms in the boron nitride structure. 6.0pt

Furthermore, B-H compounds, called boranes, are an important class of boron compounds. The simplest stable borane is diborane ( $\text{B}_2\text{H}_6$ ) and many of the higher boranes can be prepared by the pyrolysis of diborane. Diborane can be synthesized via metathesis of a boron halide and a hydride source.

**8.8** **Write** a balanced chemical equation for the synthesis of diborane from the reaction of  $\text{BF}_3$  and  $\text{LiBH}_4$ . **Hint:** both products are boron compounds. 3.0pt

**8.9** **Draw** the molecular geometry of the diborane molecule. **Hint:** there is no B-B bond in the molecule. 2.0pt

$\text{BH}_3$  (borane) is an unstable and highly reactive molecule. Therefore, it is not possible to isolate it as  $\text{BH}_3$  under ordinary conditions. However, it can be stabilized via its reaction with carbon monoxide to yield borane carbonyl ( $\text{BH}_3\text{CO}$ ) compound, which is an adduct of borane. The preparation of  $\text{BH}_3\text{CO}$  plays an important role in exploring the chemistry of boranes as it indicates the likely existence of the borane molecule.

**8.10** **Sketch** the Lewis dot structure of  $\text{BH}_3\text{CO}$  molecule by showing the formal charges. 3.0pt

**8.11** Which of the statements given in the answer sheet is observed in the C—O bond of a CO molecule upon the bond formation between  $\text{BH}_3$  and CO? **Tick** the correct box. 2.0pt

Borazine consists of single and double bonded cyclic B—N units and hydrogen atoms attached to these atoms with the molecular formula  $\text{B}_3\text{N}_3\text{H}_6$  and is isostructural to benzene. Borazine can be synthesized by using a two-step procedure including the synthesis of symmetrically trisubstituted chlorine derivatives of borazine ( $\text{B}_3\text{N}_3\text{H}_3\text{Cl}_3$ ) from the reaction of ammonium chloride and boron trichloride, and then reduction of  $\text{B}_3\text{N}_3\text{H}_3\text{Cl}_3$  with  $\text{LiBH}_4$  in THF.



**8.12** **Write** the balanced chemical equations for the two-step synthesis of borazine starting from ammonium chloride in THF (tetrahydrofuran). **Hint:** THF stabilizes one of the products by forming a Lewis acid-base adduct. 4.0pt

**8.13** **Draw** the molecular structures of borazine and its symmetrically trisubstituted chlorine derivative. 4.0pt

Catalysts are substances that accelerate the rate of reactions by allowing them to proceed in a lower energy pathway. The catalytic activity of the catalysts is generally determined by the turnover frequency (TOF), which is calculated by dividing the molar amount of the product to the mole of active catalyst and the time ( $\text{TOF} = \text{mole product} / (\text{mole catalyst} \times \text{time})$ ). A typical hydrolysis of  $\text{BNH}_6$  was carried out in 10.0 mL of water by using 100.0 mM  $\text{BNH}_6$  and 5.0 mg of CuPt/C catalyst (CuPt alloy nanoparticles supported on carbon black containing 8.2 wt% Pt atom). 67.25 mL of hydrogen gas was generated in 5 minutes.

**8.14** Assuming the catalytic reaction is performed in standard conditions (1 atm and 273.15 K), **calculate** the TOF ( $\text{min}^{-1}$ ) of the CuPt/C catalyst **in terms of only Pt atoms** in the hydrolysis of  $\text{BNH}_6$  by considering the volume of the hydrogen gas generated. 4.0pt

As a result of detailed crystal analysis of a synthesized  $\text{Cu}_x\text{Pt}_y$  alloy nanoparticle (the subscripts indicate molar percentages of the atoms in the alloy structure), it was determined that the face centered cubic (fcc) unit cell was formed by Pt atoms and the Pt atoms on the face of the fcc unit cell are supposed to be replaced with Cu atoms to form  $\text{Cu}_x\text{Pt}_y$  displacement alloy nanoparticles. According to this information, answer the following questions.

**8.15** **Determine** the composition of the alloy nanoparticles by finding x and y in the  $\text{Cu}_x\text{Pt}_y$  alloy composition. 2.0pt

**8.16** **Sketch** the shape of the described crystal unit cell of  $\text{Cu}_x\text{Pt}_y$  alloy nanoparticles by showing the placement of atoms on the unit cell. 2.0pt

**8.17** Another alloy has a  $\text{Cu}_2\text{Pt}_1$  composition. Assume that this alloy also has an fcc unit cell with an edge length of 380 pm, but the Cu and Pt atoms are randomly distributed in the atomic positions. **Calculate** the density of this alloy in  $\text{g/cm}^3$ . 4.0pt



## Boron Compounds and Hydrogen Storage

**8.1** (3.0 pt)

**8.2** (3.0 pt)

**8.3** (4.0 pt)

**8.4** (4.0 pt)



**8.5** (4.0 pt)

**8.6** (4.0 pt)

**8.7** (6.0 pt)

$\text{B}_3\text{O}_6^{3-}$	Boric acid	Boron nitride



**8.8** (3.0 pt)

**8.9** (2.0 pt)

**8.10** (3.0 pt)





**8.11** (2.0 pt)

- ☐ It gets longer because there will be  $\pi$ -back donation from  $\text{BH}_3$  to CO.
- ☐ It gets longer because CO donates  $\pi$ -bonding electrons to  $\text{BH}_3$
- ☐ No or slight change on it because CO donates its mainly non-bonding electrons to  $\text{BH}_3$
- ☐ It gets shorter because CO donates  $\pi^*$  anti-bonding electrons to  $\text{BH}_3$ .

**8.12** (4.0 pt)

**8.13** (4.0 pt)

**8.14** (4.0 pt)



**8.15** (2.0 pt)

**8.16** (2.0 pt)

**8.17** (4.0 pt)



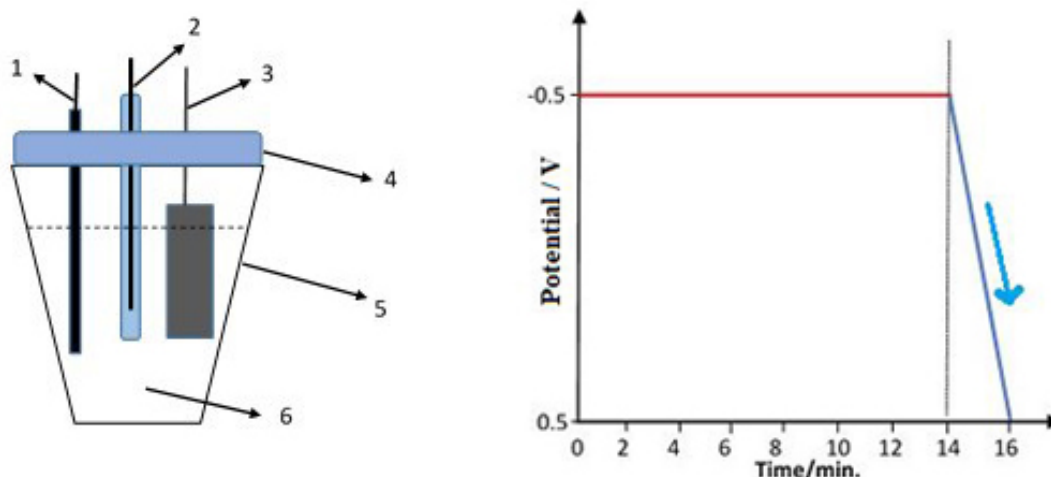
## Quantification of Heavy Metal Ions

For the quantitative analysis of heavy metal ions in a factory's wastewater pool, the following steps have been applied by an analyzer at 298 K:

**Step 1)** 10-mL samples of each were obtained from five different regions in a wastewater pool, mixed in a 100-mL beaker, and then stirred for 5 minutes using a magnetic stirrer.

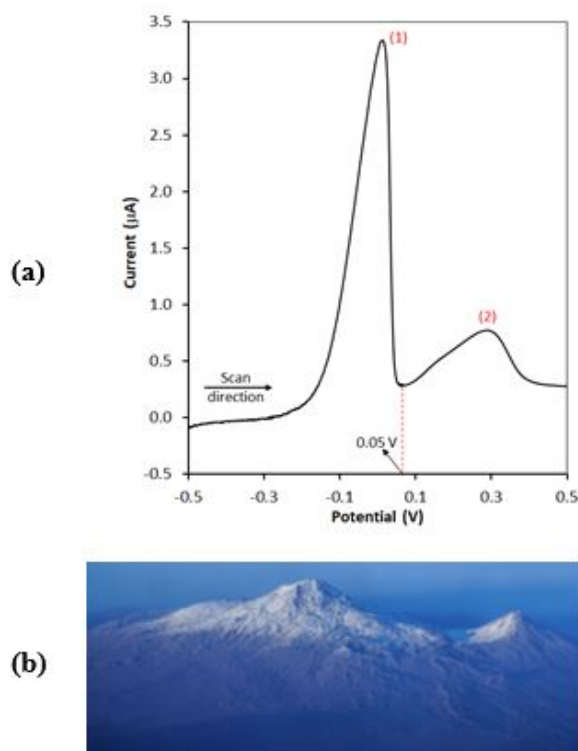
**Step 2)** 10-mL of sample solution was taken from the 100-mL beaker and 142 mg of  $\text{Na}_2\text{SO}_4$  was added while stirring, followed by transfer to a three-electrode cell as seen in Figure 1a. In this electrochemical cell, Pt wire, Ag/AgCl (3 M KCl), and Pt foil served as the working, reference, and counter electrodes, respectively.

**Step 3)** These electrodes were connected to a potentiostat and a constant potential of  $-0.50\text{ V}$  vs. Ag/AgCl for 14 minutes was applied as seen in Figure 1b (horizontal line). It is assumed that 14 min. is sufficient to complete the expected electrochemical reactions.



**Figure 1. a)** Electrochemical cell design; 1) Working electrode (Pt wire), 2) reference electrode (Ag/AgCl, 3M KCl), 3) counter electrode (Pt foil), 4) cell tap, 5) electrochemical cell, 6) 10-mL of sample solution. **b)** Potential change of working electrode as a function of time. y-axis: potential/V vs Ag/AgCl, x-axis: time/min.

**Step 4)** The electrodes were rinsed with distilled water, placed into another electrochemical cell including 10-mL of  $0.1\text{ M H}_2\text{SO}_4$  solution, and potential was scanned between  $-0.50$  and  $+0.50\text{ V}$  as seen in Figure 1b (downward sloping line in 2 min.). Current vs. potential data for this step are presented in Figure 2a, which is like an excellent view of *Mount Ararat (Ağrı Dağı)*, the highest mountain in Turkey (Figure 2b).

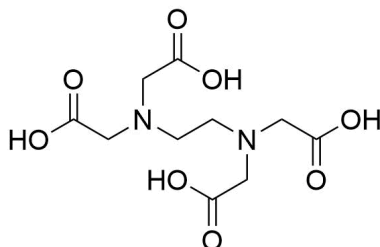


**Figure 2.** a) Potential scan of the working electrode as a function of current in 0.1 M  $\text{H}_2\text{SO}_4$  solution after keeping it constant at a potential of  $-0.50$  V in 10-mL of wastewater sample as seen in Figure 1b (horizontal line). y-axis: current/ $\mu\text{A}$ , x-axis: potential/V vs Ag/AgCl, b) A view of Great and Little Ararat peaks.

**Step 5)** Another 10-mL of the sample solution prepared in *step 1* was taken and the processes explained in *steps 2 and 3* were applied in that order. The electrodes were rinsed with distilled water and placed into 10-mL of 0.1 M  $\text{H}_2\text{SO}_4$  solution. Then the potential of the working electrode was kept constant at  $+0.05$  V for 14 min. It is assumed that 14 minutes is sufficient to complete the expected electrochemical reactions.

**Step 6)** After *step 5* was performed, the solution in the electrochemical cell was placed in a suitable oven to evaporate at  $150^\circ\text{C}$  until dry solid was obtained.

**Step 7)** 5-mL of ethylenediaminetetraacetic acid (EDTA,  $\text{H}_4\text{Y}$ ) (Figure 3) solution was added to the solid obtained in step 6 and it was shaken to dissolve. It is known that 1-mL of EDTA solution is equivalent to 3.85 mg/mL  $\text{BaCO}_3$ . Then, pH of the solution was adjusted to 10.0. Excess EDTA was titrated with standard 0.0010 M  $\text{Ni}(\text{NO}_3)_2$  solution and it was observed that 95.60 mL of  $\text{Ni}(\text{NO}_3)_2$  solution was consumed up to the endpoint.



**Figure 3.** Chemical structure of EDTA ( $H_4Y$ ).

- In water saturated with  $H_2S$ , equilibrium concentration of  $[H_2S]$  is 0.1 M.
- $K_{sp}(NiS) = 4.0 \times 10^{-20}$ ;  $K_{sp}(CuS) = 1.0 \times 10^{-36}$
- $K_{a1}(H_2S) = 9.6 \times 10^{-8}$ ;  $K_{a2}(H_2S) = 1.3 \times 10^{-14}$

Reaction	$E^\circ / V$ (at 298 K)
$2H_2O(l) + 2e^- \rightarrow H_2(g) + 2OH^-(aq)$	-0.83
$Ni^{2+}(aq) + 2e^- \rightarrow Ni(s)$	-0.24
$2H^+(aq) + 2e^- \rightarrow H_2(g)$	0.00
$Cu^{2+}(aq) + 2e^- \rightarrow Cu(s)$	+0.34
$Ag^+(aq) + e^- \rightarrow Ag(s)$	+0.80
$O_2(g) + 4H^+(aq) + 4e^- \rightarrow 2H_2O(l)$	+1.23

**9.1** Which of the following can be considered for peak 1 and peak 2 in Figure 2a, respectively? **Tick** the correct box on the answer sheet. 5.0pt

**9.2** Which of the statements is expected, if the potential is applied as  $-1.2$  V instead of  $-0.5$  V at the first step (horizontal line) in Figure 1b? **Tick** the correct box on the answer sheet. 5.0pt

**9.3** **Calculate** the scan rate of data presented in Figure 2a as mV/s at 298 K. 8.0pt

The potential of the following cell is measured as 0.437 V.

$Pt, H_2(0.92 \text{ bar}) | HCl(1.50 \times 10^{-2} M), AgCl(sat) | Ag$

**9.4** **Calculate** the standard electrode potential value (V) of half-cell of  $AgCl(s) + e^- \rightarrow Ag(s) + Cl^-(aq)$  at 298 K. 16.0pt  
Note: You must show all works.

**9.5** Which of the statements is the main purpose of step 5 in this analysis? **Tick** the correct box on the answer sheet. 5.0pt



**9.6** **Write** net ionic equations for the complexation and the back titration reaction of *step 7* on the answer sheet. 6.0pt

**9.7** **Calculate**  $\text{Ni}^{2+}$  concentration as *mg/L* in the wastewater of the factory. *Note:* You must show all works. 25.0pt

**9.8** **Calculate** the minimum pH value for starting the precipitation of  $\text{Ni}^{2+}$  ions in the solution obtained in *step 5* by passing  $\text{H}_2\text{S}$  gas into the solution until saturation. If you cannot solve question **9.7**, use 20 *mg/L*  $\text{Ni}^{2+}$  *sample* for this question. *Note:* You must show all works. 30.0pt

**Quantification of Heavy Metal Ions****9.1** (5.0 pt)

- ☐ Peak 1: electrochemical reduction of Ni / Peak 2: electrochemical reduction of Cu
- ☐ Peak 1: electrochemical reduction of Cu / Peak 2: electrochemical reduction of Ni
- ☐ Peak 1: electrochemical reduction of Ni / Peak 2: electrochemical oxidation of Cu
- ☐ Peak 1: electrochemical oxidation of Ni / Peak 2: electrochemical oxidation of Cu
- ☐ Peak 1: electrochemical oxidation of Cu / Peak 2: electrochemical oxidation of Ni

**9.2** (5.0 pt)

- ☐ NO evolution
- ☐ NO<sub>2</sub> evolution
- ☐ Nitrogen evolution
- ☐ Oxygen evolution
- ☐ Hydrogen evolution

**9.3** (8.0 pt)

Show your calculation:

Scan rate = ..... mV/s



**9.4** (16.0 pt)

Show your calculation:

Standard electrode potential = ..... V

**9.5** (5.0 pt)

- ☐ Modification of Pt wire with Ni-Cu alloy film
- ☐ Modification of Pt wire with Ni film
- ☐ Electrochemical stripping of both Cu and Ni from Cu-Ni-modified Pt wire to the solution
- ☐ Electrochemical stripping of Cu from Cu-Ni-modified Pt wire to the solution
- ☐ Electrochemical stripping of Ni from Cu-Ni-modified Pt wire to the solution

**9.6** (6.0 pt)

Complexation:

Back titration:





9.7 (25.0 pt)

Show your calculation:

Ni<sup>2+</sup> concentration: ..... mg/L:



**9.8** (30.0 pt)

Show your calculation:

Minimum pH value: .....